



## Crash Course Chemistry: Outtakes #3

### Outtakes

[https://youtube.com/watch?v=7aXoo-\\_7Pco](https://youtube.com/watch?v=7aXoo-_7Pco)

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-intro music-

-Well it turns out it's not really the universe's fault if you think about it there's only at best--nope. Well if you--if--well it turns out, well it turns out it's not really the universe's fault if you think about it there's only one way or at best maybe a few ways for things to be organized and, nope.

-Think of it this way, your lunch, was a very complicated and organized set of molecules, you ate it...and then something happened [laughs] you--it gave you the energy to --okay and well I just wanted to change this a little. It's strange we're always like assuming our audience has done something. Like that time you got your lip pierced at the mall and I was like "No! No one did that!" [laughing in background]

-Those molecules were broken down even further as they were converted--ooh. Those molecules were then broken down even further as they were converted into a dadadadadaugh!

-I gotta go back to wide.

-By the time you've finished your house maybe orbrally  
-[off camera 'nopes'].

-We all have to live with disorders so you might as--we all have to live with dist--we all have to live [horse noise]

-So even though we sometimes think of spontaneous as meaning sudden or impulsive like the majority of mall lip piercings in chemistry spontaneous I dthzatha.

-Maybe you're casually dating or in a committed relationship or maybe dablughada.

-How come male humans never fall for female vampires? It's never happened.

-It's not about drama here in chemistry, we don't want the drama. In real life or at least Paula Abdul songs. That's a Paula Abdul song right?  
-[off camera] yes! with MC Scat Cat. [singing].

-I'm Mold! I'll just keep a constant loud so you can get it ahhhhhhhh!  
[singing].

-Bonds can also be ionic if the electrons are transferred and you learned to calculate the energy of an ionic bond using [gibberish] bunny rabbit.

-Molecules!  
-[off camera] That's it! We're done!

-It's a kind of symmetry not just of the molecule but of the charge of--euh.

-Okay, now here are v-- are two v--oh right we're not doing this yet  
-[off camera] Uh-oh  
-[singing] So, okay. So- close. [laughing]

-They'll do anything just to be near it. Why? Well some simple sup-sss.. Are you serious, Hank? Are you serious? Got the whole thing in one take, and then you said 'simple seriotic sands'?

-Electronegativity increases from left to right on the period table because there are what what what what what.

-So for polarity to occur in a molecule you have to have two different elements at a minimum, and the difference between their

electronegativities has to be greater than .5

-[off camera] Let's do that again.

-Okay.

-This is a version of a molecule that's great eughh.

-But I wanna take you beyond the beginner stuff and help you understand some mode- models that explain some m-mr [mumbling]

-And to be honest this concept of electrons holding everything together is itself a model [coughs]

-[burps]

-[tongue clicking]

-[clears throat] [coughs]

-I mean think of the human models, the, that peop- okay.

-Do you think that the women in underwear catalogues and the guys whoa!

-So in addition to understanding how a model represents reality you also have ways to recognize the ways it doesn't represent reali- oh op that was, something went wrong in my head.

-So you don't base a bunch of incorrect ... assumptions on it. We can just cut, edit that pause out?

-So in addition to understanding how a model represents reality you also have to recognize the ways in what the [bleep]

-And thanks to you for watching this episode of Crash Course Chemistry, I think you're all model students  
-[off camera] Wow.

-It falls as rain, runs as rivers through beds of blaghh

-So, there we have it, gallium! By the power of my body I have turned it into a liquid and that just makes it very clear to me how whiten-- blaghh. Wick lick bow!

-These are the weakest of the inter-molecular forces because they're based on the temporary [mouths 'temporary']

-Which makes it sound like I think London is small and weak; I don't I [weird tongue noises]

-Like, I can smash ice with a hammer all I want and break it into smaller and smaller pieces, but the only way to e- but only a big enough [mutters]

-Conversely, the particles in a solid are very close together, so they have high density and can be compressed only a teeny tiny bit even under huge amounts of pressure, and I'm not even [gibberish]  
-[off camera laughing]

-Wah, wah, wah, wah. Close? For one thing [laughs] for waah  
-[off camera laughing]

-Generally a liquid can't form a perfect sphere because it's not in space [laughs]

-Cohesion manifests itself in other ways too, for example in molecules with very large inter-molecular forces like-  
-[off camera laughing]

-I was, I was gonna say honey but honey is not a molecule it is a



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bunch of honeys, it is a bunch of molecules.

-[off camera] It is a bunch of honeys.

-It is a bunch of honeys!

-Cohesion manifests itself in other ways too, for example in molecules with very large inter-molecular forces like the molecules in honey, which has a bunch of oxygen and hydrogens all wanting to hydrogen bond with each other, the cohesion is so great that it makes the liquid flow very slowly - yes I know that wasn't very good.

[laughing]

-[off camera laughing] Like, think that's okay, huh?

-That doesn't make any sense!

-And finally, you learn that surface tension can be broken by substances I didn't ev- I cut that out.

-Our script supervisor is Katherine Greend- Greend?! I know your name, you're my wife!

-[off camera] It's your name too!

-It's my name too!

-And unfortunately for all of us that brings to the end of this days episode of Crash Course Chemistray! [laughs]

-Also you learned about Henry's Law, which states that the concentration of a dissolved gas equals the partial pressure of that gas above the solution, a-[gibberish] [laughing]

-Q, you've learned how to calculate the equilibrium conditions of reactions just for, just from, just from knowing their...

-And you may even have learned a little bit about the quadratic equation [weird tongue noises] [laughing]

-outro music-